

JEE Main 29th Jan Shift 2 Question Paper

Time Allowed :3 Hours

Maximum Marks :300

Total Questions :75

General Instructions

Read the following instructions very carefully and strictly follow them:

1. This question paper contains 75 questions. All questions are compulsory.
2. This question paper is divided into three section - Mathematics, Physics and Chemistry.
3. MCQs: Four marks will be awarded for each correct answer and there will be a negative marking of one mark on each wrong question.

(Mathematics) Section-A

1. If the set of all $a \in \mathbb{R}$, for which the equation $2x^2 + (a - 5)x + 15 = 3a$ has no real root, is the interval (α, β) , and $X = \{x \in \mathbb{Z} : \alpha < x < \beta\}$, then $\sum_{x \in X} x^2$ is equal to:

- (1) 2109
- (2) 2129
- (3) 2139
- (4) 2119

2. If $\sin x + \sin^2 x = 1$, $x \in (0, \frac{\pi}{2})$, then the expression

$$(\cos^2 x + \tan^2 x) + 3(\cos^4 x + \tan^4 x + \cos^4 x + \tan^4 x) + (\cos^6 x + \tan^6 x)$$

is equal to:

- (1) 4
- (2) 3
- (3) 2
- (4) 1

3. Let the area enclosed between the curves $|y| = 1 - x^2$ and $x^2 + y^2 = 1$ be α . If

$9\alpha = \beta\pi + \gamma$; β, γ are integers, then the value of $|\beta - \gamma|$ equals:

- (1) 27
 - (2) 18
 - (3) 15
 - (4) 33
-

4. If the domain of the function $\log_5(18x - x^2 - 77)$ is (α, β) and the domain of the function

$$\log_{(x-1)} \left(\frac{2x^2 + 3x - 2}{x^2 - 3x - 4} \right)$$

is (γ, δ) , then $\alpha^2 + \beta^2 + \gamma^2$ is equal to:

- (1) 195
 - (2) 174
 - (3) 186
 - (4) 179
-

5. Let the function $f(x) = (x^2 - 1)|x^2 - ax + 2| + \cos |x|$ be not differentiable at the two points $x = \alpha = 2$ and $x = \beta$. Then the distance of the point (α, β) from the line

$12x + 5y + 10 = 0$ is equal to:

- (1) 3
 - (2) 4
 - (3) 2
 - (4) 5
-

6. Let a straight line L pass through the point $P(2, -1, 3)$ and be perpendicular to the lines

$$\frac{x-1}{2} = \frac{y+1}{1} = \frac{z-3}{-2} \quad \text{and} \quad \frac{x-3}{1} = \frac{y-2}{3} = \frac{z+2}{4}.$$

If the line L intersects the yz -plane at the point Q , then the distance between the points P and Q is:

- (1) 2
 - (2) $\sqrt{10}$
 - (3) 3
 - (4) $2\sqrt{3}$
-

7. Let $S = \mathbb{N} \cup \{0\}$. Define a relation R from S to \mathbb{R} by:

$$R = \left\{ (x, y) : \log_e y = x \log_e \left(\frac{2}{5} \right), x \in S, y \in \mathbb{R} \right\}.$$

Then, the sum of all the elements in the range of R is equal to:

- (1) $\frac{3}{2}$
 - (2) $\frac{5}{3}$
 - (3) $\frac{10}{9}$
 - (4) $\frac{5}{2}$
-

8. Let the line $x + y = 1$ meet the axes of x and y at A and B , respectively. A right-angled triangle AMN is inscribed in the triangle OAB , where O is the origin and the points M and N lie on the lines OB and AB , respectively. If the area of the triangle AMN is $\frac{4}{9}$ of the area of the triangle OAB and $AN : NB = \lambda : 1$, then the sum of all possible values of λ is:

- (1) $\frac{1}{2}$
 - (2) $\frac{13}{6}$
 - (3) 5
 - (4) 2
-

9. If $\alpha x + \beta y = 109$ is the equation of the chord of the ellipse

$$\frac{x^2}{9} + \frac{y^2}{4} = 1$$

whose midpoint is $\left(\frac{5}{2}, \frac{1}{2}\right)$, then $\alpha + \beta$ is equal to:

- (1) 37
- (2) 46
- (3) 58

(4) 72

10. If all the words with or without meaning made using all the letters of the word "KANPUR" are arranged as in a dictionary, then the word at 440th position in this arrangement is:

- (1) PRNAKU
- (2) PRKANU
- (3) PRKAUN
- (4) PRNAUK

11. Let $\alpha, \beta (\alpha \neq \beta)$ be the values of m , for which the equations $x + y + z = 1$, $x + 2y + 4z = m$, and $x + 4y + 10z = m^2$ have infinitely many solutions. Then the value of $\sum_{n=1}^{10} (n^4 + n^8)$ is equal to:

- (1) 440
- (2) 3080
- (3) 3410
- (4) 560

12. Let $A = [a_{ij}]$ be a matrix of order 3×3 , with $a_{ij} = (\sqrt{2})^{i+j}$. If the sum of all the elements in the third row of A^2 is $\alpha + \beta\sqrt{2}$, where $\alpha, \beta \in \mathbb{Z}$, then $\alpha + \beta$ is equal to:

- (1) 280
- (2) 168
- (3) 210
- (4) 224

13. Let P be the foot of the perpendicular from the point $(1, 2, 2)$ on the line

$$\frac{x-1}{1} = \frac{y+1}{-1} = \frac{z-2}{2}$$

Let the line $\mathbf{r} = (-\hat{i} + \hat{j} - 2\hat{k}) + \lambda(\hat{i} - \hat{j} + \hat{k})$, $\lambda \in \mathbb{R}$, intersect the line L at Q . Then $2(PQ)^2$ is equal to:

- (1) 27

- (2) 25
 - (3) 29
 - (4) 19
-

14. Let a circle C pass through the points (4, 2) and (0, 2), and its centre lie on $3x + 2y + 2 = 0$. Then the length of the chord of the circle C, whose midpoint is (1, 2), is:

- (1) $\sqrt{3}$
 - (2) $2\sqrt{3}$
 - (3) $4\sqrt{2}$
 - (4) $2\sqrt{2}$
-

15. Let $A = [a_{ij}]$ be a 2×2 matrix such that $a_{ij} \in \{0, 1\}$ for all i and j . Let the random variable X denote the possible values of the determinant of the matrix A. Then, the variance of X is:

- (1) $\frac{1}{4}$
 - (2) $\frac{3}{8}$
 - (3) $\frac{5}{8}$
 - (4) $\frac{3}{4}$
-

16. Bag 1 contains 4 white balls and 5 black balls, and Bag 2 contains n white balls and 3 black balls. One ball is drawn randomly from Bag 1 and transferred to Bag 2. A ball is then drawn randomly from Bag 2. If the probability that the ball drawn is white is $\frac{29}{45}$, then n is equal to:

- (1) 3
 - (2) 4
 - (3) 5
 - (4) 6
-

17. The remainder, when 7^{98} is divided by 23, is equal to:

- (1) 14
- (2) 9

(3) 17

(4) 6

18. Let $f(x) = \int_1^x (t^2 - 9t + 20) dt$, $1 \leq x \leq 5$. If the range of $f(x)$ is $[\alpha, \beta]$, then $4(\alpha + \beta)$ equals:

(1) 157

(2) 253

(3) 125

(4) 154

19. Let \hat{a} be a unit vector perpendicular to the vectors

$$\mathbf{b} = \hat{i} - 2\hat{j} + 3\hat{k} \quad \text{and} \quad \mathbf{c} = 2\hat{i} + 3\hat{j} - \hat{k},$$

and makes an angle of $\cos\left(-\frac{1}{3}\right)$ with the vector $\hat{i} + \alpha\hat{j} + \hat{k}$. If \hat{a} makes an angle of $\frac{\pi}{3}$ with the vector $\hat{i} + \alpha\hat{j} + \hat{k}$, then the value of α is:

(1) $\sqrt{3}$

(2) $\sqrt{6}$

(3) $\sqrt{3}$

(4) $\sqrt{6}$

20. If for the solution curve $y = f(x)$ of the differential equation

$$\frac{dy}{dx} + (\tan x)y = 2 + \sec^2 x, \quad y\left(\frac{\pi}{3}\right) = \sqrt{3},$$

then $y\left(\frac{\pi}{4}\right)$ is equal to:

(1) $\frac{3+\sqrt{3}}{2}$

(2) $\frac{3+1}{(1+\sqrt{3})}$

(3) $\frac{3+\sqrt{3}}{(4+\sqrt{3})}$

(4) $\frac{4-\sqrt{2}}{14}$

(Mathematics) Section-B

21. If

$$24 \left(\int_0^{\frac{\pi}{4}} \left[\sin \left(4x - \frac{\pi}{12} \right) + [2 \sin x] \right] dx \right) = 2n + \alpha,$$

where $[.]$ denotes the greatest integer function, then α is equal to:

22. If $\lim_{t \rightarrow \infty} \left(\int_0^1 (3x + 5)^t dx \right) = \frac{\alpha}{5e} \left(\frac{8}{5} \right)^{\frac{3}{2}}$, then α is equal to ---- :

23. Let $a_1, a_2, \dots, a_{2024}$ **be an Arithmetic Progression such that**

$$a_1 + (a_1 + a_0 + a_1 + a_2 + \dots + a_{2020} + a_{2024}) = 2233. \quad \text{Then} \quad a_1 + a_2 + a_3 + \dots + a_{2022}$$

is equal to ---- :

24. Let integers $a, b \in [-3, 3]$ be such that $a + b \neq 0$.

Then the number of all possible ordered pairs (a, b) , for which

$$\left| \frac{z-a}{z+b} \right| = 1 \quad \text{and} \quad \begin{vmatrix} z+1 & \omega & \omega^2 \\ \omega^2 & 1 & z+\omega \\ \omega^2 & 1 & z+\omega \end{vmatrix} = 1,$$

is equal to:

25. Let $y^2 = 12x$ **be the parabola and S its focus. Let PQ**

be a focal chord of the parabola such that $(SP)(SQ) = \frac{147}{4}$.

Let C be the circle described taking PQ as a diameter. If the equation of a circle C is

$64x^2 + 64y^2 - \alpha x - 64\sqrt{3}y = \beta$, then $\beta - \alpha$ is equal to:

(Physics) Section-A

26. The difference of temperature in a material can convert heat energy into electrical energy. To harvest the heat energy, the material should have:

(1) low thermal conductivity and low electrical conductivity

- (2) high thermal conductivity and high electrical conductivity
 - (3) low thermal conductivity and high electrical conductivity
 - (4) high thermal conductivity and low electrical conductivity
-

27. Given below are two statements. One is labelled as Assertion (A) and the other is labelled as Reason (R).

Assertion (A): With the increase in the pressure of an ideal gas, the volume falls off more rapidly in an isothermal process in comparison to the adiabatic process.

Reason (R): In an isothermal process, $PV = \text{constant}$, while in an adiabatic process $PV^\gamma = \text{constant}$. Here, γ is the ratio of specific heats, P is the pressure and V is the volume of the ideal gas.

In the light of the above statements, choose the correct answer from the options given below:

- (1) Both (A) and (R) are true but (R) is NOT the correct explanation of (A)
 - (2) (A) is true but (R) is false
 - (3) Both (A) and (R) are true and (R) is the correct explanation of (A)
 - (4) (A) is false but (R) is true
-

28. An electric dipole is placed at a distance of 2 cm from an infinite plane sheet having positive charge density σ . Choose the correct option from the following.

- (1) Torque on dipole is zero and net force is directed away from the sheet.
 - (2) Torque on dipole is zero and net force acts towards the sheet.
 - (3) Potential energy of dipole is minimum and torque is zero.
 - (4) Potential energy and torque both are maximum.
-

29. In an experiment with photoelectric effect, the stopping potential:

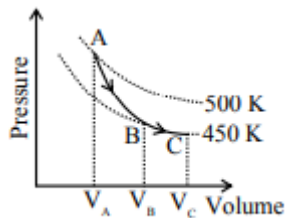
- (1) increases with increase in the wavelength of the incident light
 - (2) increases with increase in the intensity of the incident light
 - (3) is $\frac{1}{e}$ times the maximum kinetic energy of the emitted photoelectrons
 - (4) decreases with increase in the intensity of the incident light
-

30. A point charge causes an electric flux of $-2 \times 10^4 \text{ Nm}^2 \text{ C}^{-1}$ to pass through a

spherical Gaussian surface of 8.0 cm radius, centered on the charge. The value of the point charge is:

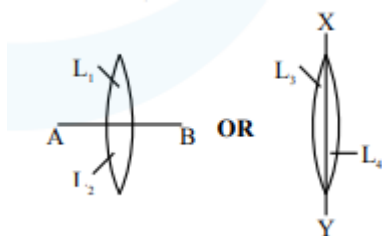
- (1) $17.7 \times 10^{-7} \text{ C}$
- (2) $15.7 \times 10^{-7} \text{ C}$
- (3) $17.7 \times 10^{-6} \text{ C}$
- (4) $15.7 \times 10^{-6} \text{ C}$

31. A poly-atomic molecule (C_3R , $C_v = 4R$, where R is gas constant) goes from phase space point A ($P_A = 10^4 \text{ Pa}$, $V_A = 4 \times 10^{-3} \text{ m}^3$) to point B ($P_B = 5 \times 10^4 \text{ Pa}$, $V_B = 6 \times 10^{-7} \text{ m}^3$) to point C ($P_C = 10^4 \text{ Pa}$, $V_C = 8 \times 10^{-3} \text{ m}^3$). A to B is an adiabatic path and B to C is an isothermal path. The net heat absorbed per unit mole by the system is:



- (1) $500R(\ln 3 + \ln 4)$
- (2) $450R(\ln 3)$
- (3) $500R(\ln 2)$
- (4) $400R \ln 2$

32. Two identical symmetric double convex lenses of focal length f are cut into two equal parts L_1, L_2 by the AB plane and L_3, L_4 by the XY plane as shown in the figure respectively. The ratio of focal lengths of lenses L_1 and L_3 is:



- (1) 1 : 4
- (2) 1 : 1

(3) 2 : 1

(4) 1 : 2

33. A plane electromagnetic wave propagates along the $+x$ direction in free space. The components of the electric field \vec{E} and magnetic field \vec{B} vectors associated with the wave in Cartesian frame are:

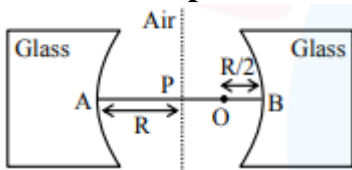
(1) E_x, B_x

(2) E_y, B_z

(3) E_z, B_y

(4) E_x, B_y

34. Two concave refracting surfaces of equal radii of curvature and refractive index 1.5 face each other in air as shown in figure. A point object O is placed midway, between P and B. The separation between the images of O, formed by each refracting surface is:



(1) 0.214R

(2) 0.114R

(3) 0.411R

(4) 0.124R

35. Two bodies A and B of equal mass are suspended from two massless springs of spring constant k_1 and k_2 , respectively. If the bodies oscillate vertically such that their amplitudes are equal, the ratio of the maximum velocity of A to the maximum velocity of B is:

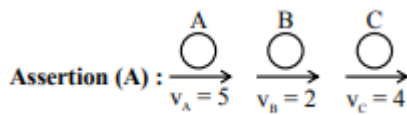
(1) $\frac{k_1}{k_2}$

(2) $\frac{k_2}{k_1}$

(3) $\sqrt{\frac{k_2}{k_1}}$

(4) $\sqrt{\frac{k_1}{k_2}}$

36. Given below are two statements, one is labelled as Assertion (A) and the other is labelled as Reason (R).



Three identical spheres of the same mass undergo one-dimensional motion as shown in the figure with initial velocities $v_A = 5 \text{ m/s}$, $v_B = 2 \text{ m/s}$, $v_C = 4 \text{ m/s}$. If we wait sufficiently long for elastic collision to happen, then $v_A = 4 \text{ m/s}$, $v_B = 2 \text{ m/s}$, $v_C = 5 \text{ m/s}$ will be the final velocities.

Reason (R): In an elastic collision between identical masses, two objects exchange their velocities.

In light of the above statements, choose the correct answer from the options given below:

- (1) Both (A) and (R) are true but (R) is NOT the correct explanation of (A)
 - (2) (A) is true but (R) is false
 - (3) Both (A) and (R) are true and (R) is the correct explanation of (A)
 - (4) (A) is false but (R) is true
-

37. A sand dropper drops sand of mass $m(t)$ on a conveyor belt at a rate proportional to the square root of the speed v of the belt, i.e., $\frac{dm}{dt} \propto \sqrt{v}$. If P is the power delivered to run the belt at constant speed, then which of the following relationships is true?

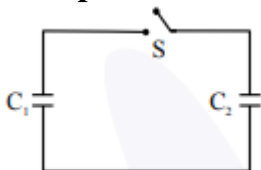
- (1) $P \propto v^3$
 - (2) $P \propto \sqrt{v}$
 - (3) $P \propto v$
 - (4) $P \propto v^5$
-

38. A convex lens made of glass (refractive index = 1.5) has a focal length of 24 cm in air. When it is totally immersed in water (refractive index = 1.33), its focal length changes to:

- (1) 72 cm
- (2) 96 cm

- (3) 24 cm
(4) 48 cm

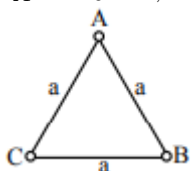
39. A capacitor, $C_1 = 6 \mu F$, is charged to a potential difference of $V_1 = 5 \text{ V}$ using a 5V battery. The battery is removed and another capacitor, $C_2 = 12 \mu F$, is inserted in place of the battery. When the switch 'S' is closed, the charge flows between the capacitors for some time until equilibrium condition is reached. What are the charges q_1 and q_2 on the capacitors C_1 and C_2 when equilibrium condition is reached?



- (1) $q_1 = 15 \mu C$, $q_2 = 30 \mu C$
(2) $q_1 = 30 \mu C$, $q_2 = 15 \mu C$
(3) $q_1 = 10 \mu C$, $q_2 = 20 \mu C$
(4) $q_1 = 20 \mu C$, $q_2 = 10 \mu C$

40. Three equal masses m are kept at vertices (A, B, C) of an equilateral triangle of side a in free space. At $t = 0$, they are given an initial velocity

$$\vec{V}_A = V_0 \hat{AC}, \vec{V}_B = V_0 \hat{BA}, \vec{V}_C = V_0 \hat{CB}.$$



Here, \hat{AC} , \hat{CB} , \hat{BA} are unit vectors along the edges of the triangle. If the three masses interact gravitationally, then the magnitude of the net angular momentum of the system at the point of collision is:

- (1) $\frac{1}{2} amv_0$
(2) $3 amv_0$
(3) $\frac{\sqrt{3}}{2} amv_0$
(4) $\frac{3}{2} amv_0$

41. Match List-I with List-II.

	List-I		List-II
(A)	Young's Modulus	(I)	$ML^{-1}T^{-1}$
(B)	Torque	(II)	$ML^{-1}T^{-2}$
(C)	Coefficient of Viscosity	(III)	$M^{-1}L^3T^{-2}$
(D)	Gravitational Constant	(IV)	ML^2T^{-2}

Choose the correct answer from the options given below:

- (1) (A)-(I), (B)-(III), (C)-(I), (D)-(II)
 - (2) (A)-(II), (B)-(IV), (C)-(II), (D)-(III)
 - (3) (A)-(III), (B)-(I), (C)-(IV), (D)-(II)
 - (4) (A)-(I), (B)-(II), (C)-(I), (D)-(IV)
-

42. Match List-I with List-II.

	List-I		List-II
(A)	Magnetic induction	(I)	Ampere meter ²
(B)	Magnetic intensity	(II)	Weber
(C)	Magnetic flux	(III)	Gauss
(D)	Magnetic moment	(IV)	Ampere meter

Choose the correct answer from the options given below:

- (1) (A)-(III), (B)-(IV), (C)-(I), (D)-(II)
 - (2) (A)-(III), (B)-(IV), (C)-(II), (D)-(I)
 - (3) (A)-(I), (B)-(IV), (C)-(III), (D)-(II)
 - (4) (A)-(IV), (B)-(III), (C)-(II), (D)-(I)
-

43. The truth table for the circuit given below is:



Choose the correct answer from the options given below:

A	B	Y
0	0	0
0	1	1
1	0	1
1	1	0

- (1)

- (2)

A	B	Y
0	0	0
1	0	0
1	1	0
0	1	1
- (3)

A	B	Y
0	0	0
1	0	1
0	1	0
1	1	0
- (4)

A	B	Y
0	0	0
1	1	1
1	0	1
0	1	1

44. A cup of coffee cools from 90°C to 80°C in t minutes when the room temperature is 20°C . The time taken by the similar cup of coffee to cool from 80°C to 60°C at the same room temperature is:

- (1) $\frac{13}{5}t$
 (2) $\frac{10}{13}t$
 (3) $\frac{13}{10}t$
 (4) $\frac{5}{13}t$

45. The number of spectral lines emitted by atomic hydrogen that is in the 4th energy level is:

- (1) 6
 (2) 3
 (3) 4
 (4) 1

(Physics) Section-B

46. The magnetic field inside a 200 turns solenoid of radius 10 cm is 2.9×10^{-4} Tesla. If the solenoid carries a current of 0.29 A, then the length of the solenoid is:

47. A parallel plate capacitor consisting of two circular plates of radius 10 cm is being charged by a constant current of 0.15 A. If the rate of change of potential difference

between the plates is 7×10^6 V/s, then the integer value of the distance between the parallel plates is:

48. A physical quantity Q is related to four observables a , b , c , and d as follows:

$$Q = \frac{ab^4}{cd^2}$$

Where: - $a = (60 \pm 3)$ Pa, - $b = (20 \pm 0.1)$ m, - $c = (40 \pm 0.2)$ N·s/m², - $d = (50 \pm 0.1)$ m.

Then the percentage error in Q is:

49. Two planets, A and B are orbiting a common star in circular orbits of radii R_A and R_B , respectively, with $R_B = 2R_A$. The planet B is $\sqrt{2}$ times more massive than planet A. The ratio $\frac{L_B}{L_A}$ of angular momentum (L) of planet B to that of planet A (L_A) is closest to integer:

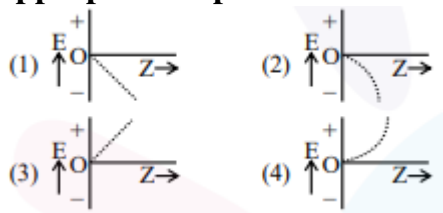
50. Two cars P and Q are moving on a road in the same direction. Acceleration of car P increases linearly with time whereas car Q moves with a constant acceleration. Both cars cross each other at time $t = 0$, for the first time. The maximum possible number of crossing(s) (including the crossing at $t = 0$) is:

(Chemistry) Section-A

51. The calculated spin-only magnetic moments of $K_3[Fe(OH)_6]$ and $K_4[Fe(OH)_6]$ respectively are:

- (1) 4.90 and 4.90 B.M.
 - (2) 5.92 and 4.90 B.M.
 - (3) 3.87 and 4.90 B.M.
 - (4) 4.90 and 5.92 B.M.
-

52. For hydrogen-like species, which of the following graphs provides the most appropriate representation of E vs Z plot for a constant n ?



53. Given below are two statements:

Statement (I): In partition chromatography, the stationary phase is a thin film of liquid present in the inert support.

Statement (II): In paper chromatography, the material of paper acts as a stationary phase.

In light of the above statements, choose the correct answer from the options given below:

- (1) Both Statement I and Statement II are false
- (2) Statement I is true but Statement II is false
- (3) Both Statement I and Statement II are true
- (4) Statement I is false but Statement II is true

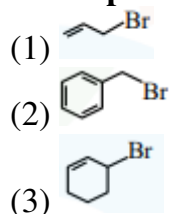
54. Identify the essential amino acids from below:

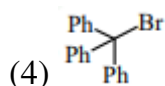
(A) Valine (B) Proline (C) Lysine (D) Threonine (E) Tyrosine

Choose the correct answer from the options given below:

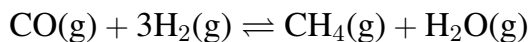
- (1) (A), (C) and (D) only
- (2) (A), (C) and (E) only
- (3) (B), (C) and (E) only
- (4) (C), (D) and (E) only

55. Which among the following halides will generate the most stable carbocation in a nucleophilic substitution reaction?





56. Consider the equilibrium:



If the pressure applied over the system increases by two fold at constant temperature then:

- (1) (A) and (B) only
 - (2) (A), (B) and (D) only
 - (3) (B) and (C) only
 - (4) (A), (B) and (C) only
-

57. Given below are two statements:

Statement (I): NaCl is added to the ice at 0°C , present in the ice cream box to prevent the melting of ice cream.

Statement (II): On addition of NaCl to ice at 0°C , there is a depression in freezing point.

In the light of the above statements, choose the correct answer from the options given below:

- (1) Statement I is false but Statement II is true
 - (2) Both Statement I and Statement II are true
 - (3) Statement I is false but Statement II is true
 - (4) Statement I is true but Statement II is false
-

58. Given below are two statements:

Statement (I): On nitration of m-xylene with HNO_3 , H_2SO_4 , followed by oxidation, 4-nitrobenzene-1, 3-dicarboxylic acid is obtained as the major product.

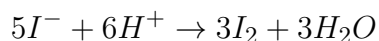
Statement (II): CH_3 group is o/p-directing while NO_2 group is m-directing group.

In the light of the above statements, choose the correct answer from the options given below:

- (1) Statement I is false but Statement II is true
- (2) Statement I is false but Statement II is true
- (3) Both Statement I and Statement II are true

(4) Statement I is true but Statement II is false

59. 0.1 M solution of KI reacts with excess of H_2SO_4 and KIO_3 , according to the equation:



Identify the correct statements: (A) 200 mL of KI solution reacts with 0.004 mol of KIO_3 (B) 200 mL of KI solution reacts with 0.006 mol of H_2SO_4 (C) 0.5 L of KI solution produced 0.005 mol of I_2 (D) Equivalent weight of KIO_3 is equal to:

$$\frac{\text{Molecular weight}}{5}$$

Choose the correct answer from the options given below:

- (1) (A) and (D) only
- (2) (A) and (B) only
- (3) (B) and (C) only
- (4) (C) and (D) only

60. Match List-I with List-II:

List-I Applications		List-II Batteries/Cell	
(A)	Transistors	(I)	Anode - Zn/Hg ; Cathode - $HgO + C$
(B)	Hearing aids	(II)	Hydrogen fuel cell
(C)	Invertors	(III)	Anode - Zn; Cathode - Carbon
(D)	Apollo space ship	(IV)	Anode - Pb ; Cathode - Pb PbO_2

Choose the correct answer from the options given below:

- (1) (A)-(III), (B)-(I), (C)-(IV), (D)-(II)
- (2) (A)-(III), (B)-(II), (C)-(I), (D)-(IV)
- (3) (A)-(II), (B)-(III), (C)-(IV), (D)-(I)
- (4) (A)-(III), (B)-(III), (C)-(IV), (D)-(I)

61. O_2 gas will be evolved as a product of electrolysis of:

- (A) an aqueous solution of AgNO_3 using silver electrodes.
(B) an aqueous solution of AgNO_3 using platinum electrodes.
(C) a dilute solution of H_2SO_4 using platinum electrodes.
(D) a high concentration solution of H_2SO_4 using platinum electrodes.

Choose the correct answer from the options given below:

- (1) (B) and (C) only
(2) (A) and (C) only
(3) (B) and (D) only
(4) (A) and (C) only
-

62. Identify the homoleptic complexes with odd number of d-electrons in the central metal:

- (A) $[\text{FeO}_4]^{2-}$
(B) $[\text{Fe}(\text{CN})_6]^{3-}$
(C) $[\text{Fe}(\text{CN})_6]^{2-}$
(D) $[\text{CoCl}_4]^{2-}$
(E) $[\text{Co}(\text{H}_2\text{O})_6]^{3+}$

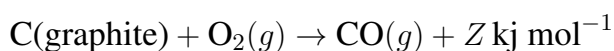
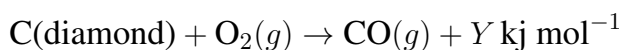
Choose the correct answer from the options given below:

- (1) (B) and (D) only
(2) (A), (B) and (D) only
(3) (A), (B) and (E) only
(4) (A), (C), (D) and (E) only
-

63. Total number of sigma (σ) and pi (π) bonds respectively present in hex-1-en-4-yne are:

- (1) 13 and 3
(2) 11 and 3
(3) 13 and 13
(4) 14 and 3
-

64. If



At constant temperature. Then:

- (1) $X = Y + Z$
 - (2) $X - Y = Z$
 - (3) $X = Y - Z$
 - (4) $X = Y + Z$
-

65. Given below are two statements:

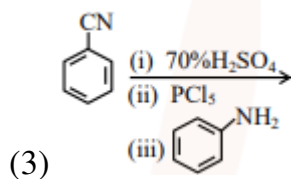
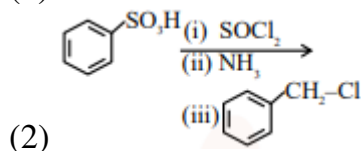
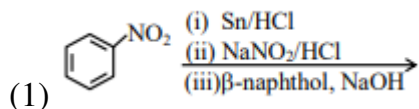
Statement (I): It is impossible to specify simultaneously with arbitrary precision, the linear momentum and the position of a particle.

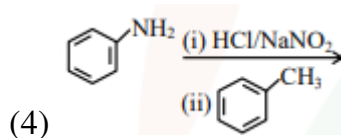
Statement (II): If the uncertainty in the measurement of position and uncertainty in measurement of momentum are equal for an electron, then the uncertainty in the measurement of velocity is $\geq \sqrt{\frac{h}{\pi}} \times \frac{1}{2m}$.

In the light of the above statements, choose the correct answer from the options given below:

- (1) Statement I is true but Statement II is false.
 - (2) Both Statement I and Statement II are true.
 - (3) Statement I is false but Statement II is true.
 - (4) Both Statement I and Statement II are false.
-

66. Which one of the following reaction sequences will give an azo dye?





67. Drug X becomes ineffective after 50% decomposition. The original concentration of drug in a bottle was 16 mg/mL which becomes 4 mg/mL in 12 months. The expiry time of the drug in months is ____ . Assume that the decomposition of the drug follows first order kinetics.

- (1) 12
- (2) 2
- (3) 3
- (4) 6

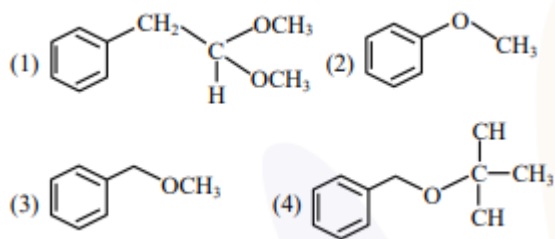
68. The type of oxide formed by the element among Li, Na, Be, Mg, B and Al that has the least atomic radius is:

- (1) A_2O_3
- (2) AO_2
- (3) A_2O
- (4) A_2O_4

69. First ionisation enthalpy values of the first four group 15 elements are given below. Choose the correct value for the element that is a main component of the apatite family:

- (1) 1012 kJ mol^{-1}
- (2) 1402 kJ mol^{-1}
- (3) 834 kJ mol^{-1}
- (4) 947 kJ mol^{-1}

70. Which one of the following, with HBr, will give a phenol?



(Chemistry) Section-B

71. Consider the following low-spin complexes



The sum of the spin-only magnetic moment values of complexes having yellow colour is:

B.M. (answer is nearest integer)

72. Isomeric hydrocarbons \rightarrow negative Baeyer's test (Molecular formula C_9H_{12}). The total number of isomers from above with four different non-aliphatic substitution sites is -

73. In the Claisen-Schmidt reaction to prepare dibenzalacetone from 5.3 g benzaldehyde, a total of 3.51 g of product was obtained. The percentage yield in this reaction was

74. In the sulphur estimation, 0.20 g of a pure organic compound gave 0.40 g of barium sulphate. The percentage of sulphur in the compound is

75. Total number of non-bonded electrons present in NO_2^- ion based on Lewis theory is:
